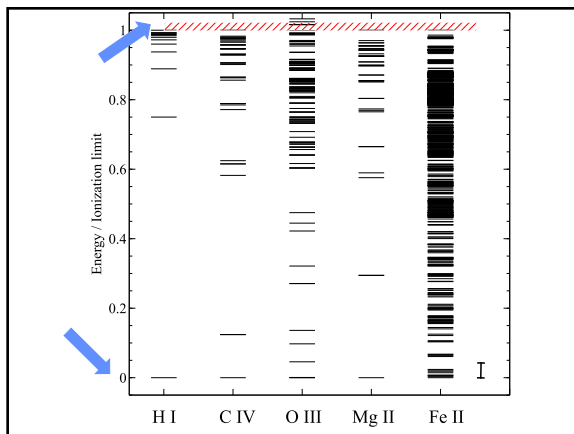
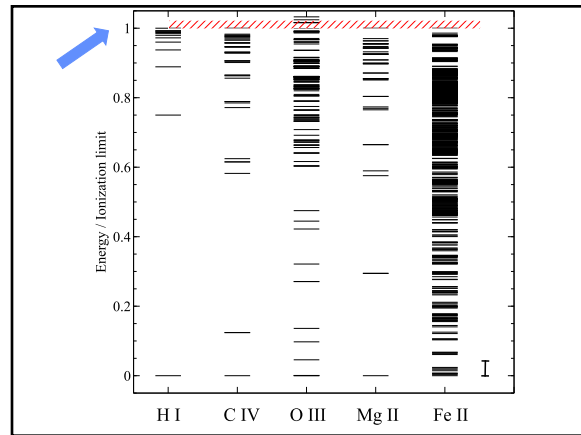
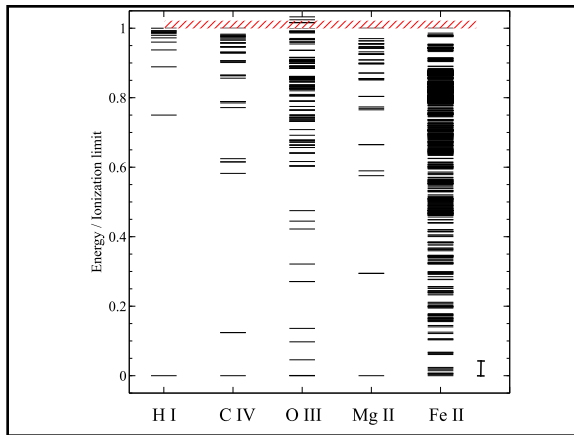
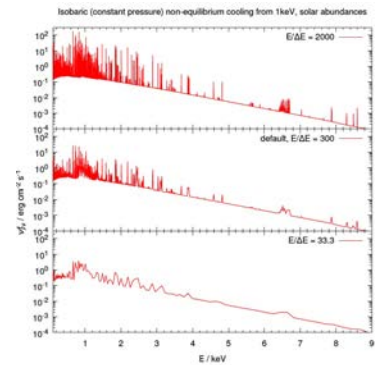


Vary blackbody temperature

- ◆ Stoy or “energy balance” method of determining stellar temperatures
- ◆ AGN3 Section 5.10

◆ How to make sense of all these lines



Peter’s atomic line list

- ◆ <http://www.pa.uky.edu/~peter/atomic/>
- ◆ <http://www.pa.uky.edu/~peter/newpage/>
– Beta version with new features
- ◆ Search wavelength range to find what lines are present

NIST

◆ <http://www.nist.gov/pml/data/asd.cfm>

Physical Measurement Laboratory

NIST Atomic Spectra Database

Version 4

Welcome to the NIST Atomic Spectra Database, NIST Standard Reference Database #78. The spectroscopic data may be selected and displayed according to wavelengths or energy levels by choosing one of the following options:

LINES Spectral lines and associated energy levels displayed in wavelength order with all selected spectra intermixed or in multiplet order. Transition probabilities for the lines are also displayed where available.

Levels Energy levels of a particular atom or ion displayed in order of energy above the ground state.

NIST ASD Team
Principal Developers (Currently Active):
Yu. Ralchenko, A. Kramida, and J. Reader

2014 Cloudy workshop

NIST Atomic Spectra Database Levels Form

Best viewed with the latest versions of Web browsers and Java

This form provides access to NIST critically evaluated data on atomic energy levels.

Spectrum: e.g., Fe I

Default Values

Level Units: Extended Search: for all levels seen

Format output:

Display output:

Page size:

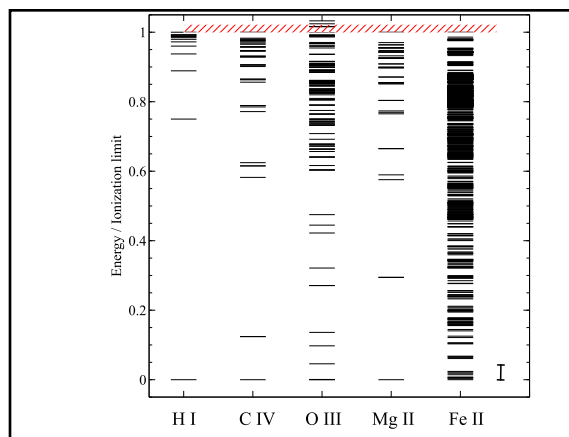
Term ordered: term energy

Energy ordered:

Level information: Principal configuration Principal term Level J Lande-g

O III

Configuration	Term	J	Level (cm ⁻¹)
2s ² 2p ²	3P	0	0
		1	113.178
		2	306.174
2s ² 2p ²	1D	2	20 273.27
2s ² 2p ²	1S	0	43 185.74
2s2p ³	5S ^o	2	60 324.79
2s2p ³	3D ^o	3	120 025.2
		2	120 053.4
		1	120 059.2



Two cases

- ◆ One and two electron systems
- ◆ Levels are closer to the continuum above than the ground state below

Two cases

- ◆ Many electron systems
- ◆ Levels have range of energies
- ◆ Many are close to the ground state below

Two types of lines

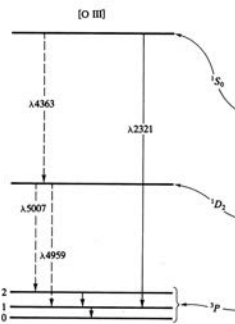
- ◆ **Recombination AGN3 sec 4.2**
 - e + p radiative recombination
 - Rate coefficient $q \sim 10^{-13} \text{ cm}^3 \text{ s}^{-1}$
 - Mainly H, He
- ◆ **Collisionally excited AGN3 3.5**
 - Inelastic e + ion collision
 - $q \sim 10^{-9} \text{ cm}^3 \text{ s}^{-1}$
 - Heavy elements

Selection rules for transitions

- ◆ AGN3
- ◆ Appendix 4 Nebular quantum mechanics
- ◆ Appendix 6 Molecular quantum

O III

Configuration	Term	J	Level (cm ⁻¹)
2s ² 2p ²	3P	0	0
		1	113.178
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2s2p ³	5S ⁺	2	60 324.79
2s2p ³	3D ⁺	3	120 025.2
		2	120 053.4
		1	120 059.2



Species vs spectra

- ◆ H⁰, C³⁺, O²⁺, H₂, CO are baryons
- ◆ H I, C IV, O III, H₂, and CO are the spectra they emit / absorb
- ◆ O III is a permitted line produced by O²⁺, while [O III] is a forbidden line
- ◆ C III] is a semi-forbidden line, often an intercombination line

Species vs spectra

- ◆ H I Ly α emission can be produced by
 - Recombination of H⁺
 - Impact excitation of H⁰
- ◆ H I absorption can only be produced by H⁰
- ◆ H I is not the same as H⁰
 - Ambiguous for emission lines

Baryons and spectra

- ◆ Hazy 1 Section 2.5
- ◆ SpeciesLabels.txt in docs
- ◆ Molecules are not ambiguous
 - H₂
 - CO
 - O₂
 - H₂⁺
 - C₂⁺
 - Their spectra have the same notation as the baryon

Baryons and spectra

- ◆ Atomic spectra use number of spectra
 - H 1
 - C 4
- ◆ The baryon
 - “H”
 - “He+”
 - “C+2”
 - (C2+ is C₂⁺ in our notation)

Lines in the main output

- ◆ Print lines column
- ◆ Print lines sort wavelength
- ◆ Print lines faint

Finding lines in Cloudy

- ◆ A line is identified by a spectral label & wavelength
- ◆ docs/LineLabels.txt has label, wavelength, comment about line
 - Generated with command “Save line labels”
- ◆ Pick lines from this file

Air vs vacuum wavelengths

- ◆ The rule in atomic physics has been to use vacuum wavelengths for $\lambda < 2000\text{\AA}$ and air for $\lambda > 2000\text{\AA}$
- ◆ SDSS has used vacuum for all wavelengths
- ◆ Today’s papers use a mix of both
- ◆ Vacuum is probably the future
- ◆ Print line vacuum
 - But you need to change your wavelengths

Some familiar lines

Species	$\lambda(\text{air})$	$\lambda(\text{vacuum})$
H 1	1215.67Å	1215.67Å
O 2	3726.03Å	3727.09Å
O 2	3728.81Å	3729.88Å
O. 3	4363.21Å	4364.44Å
H 1	4861.33Å	4862.69Å
O 3	5006.84Å	5008.24Å
H. 1	6562.81Å	6564.62Å

Other database reporting options

- ◆ See C17 review article, section 2
- ◆ Database print

Line blends

- ◆ Blnd 3727
- ◆ Blnd 2798
- ◆ Blnd 1549
- ◆ Two or more lines that appear as a single line in most spectra

Luminosity, relative intensity

- ◆ Intensity or luminosity of line
 - depending on case
- ◆ Intensity relative to normalization line, default H β

– Change with
normalize
command

0	3	88.3323m	-5.577	1.5126
0	3	51.8004m	-5.106	4.4704
0	3	4931.23A	-8.339	0.0026
0	3	4958.91A	-4.876	7.5973
0	3	5006.84A	-4.401	22.6702
0	3	2320.95A	-7.193	0.0366
0	3	4363.21A	-6.593	0.1456
0	3	1660.81A	-7.187	0.0371
0	3	1666.15A	-6.720	0.1087

Why use the laser at all

- ◆ Cloudy has lots of lines and does many levels for many ions
- ◆ A single zone (which we do for speed) is optically thin
- ◆ So continuum fluorescent excitation can be important.
- ◆ But would not happen with a finite column density
- ◆ Show fig with energy levels for H, C IV etc and say continuum photons would excite to all upper levels

Two level atom AGN3 Sec 3.5

- ◆ Excitation, deexcitation rates
- ◆ Transition probabilities
- ◆ Critical density
- ◆ Two limits
 - Low densities, every excitation leads to emission of a photon
 - high densities, levels are n LTE, photon emission proportional to $n_u A_{ul}$

$$4\pi j = n_u A_{ul} h\nu$$

$$[\text{erg cm}^{-3} \text{s}^{-1}]$$

$$n_e g_{lu} n_e = n_u [A_{ul} + g_{ue} n_e]$$

$$\frac{n_u}{n_e} = \frac{g_{lu} n_e}{A_{ul} + g_{ue} n_e}$$

$$n_u + n_e = n$$

critical density

$$A_{ul} = g_{ue} n_{crit}$$

LDL

$$n_e \ll n_{crit}$$

$$4\pi j = n_e n_e g_{lu} h\nu$$

HDL

$$n_e \gg n_{crit}$$

$$4\pi j = n_e \frac{g_{lu} A_{ul} h\nu}{g_{ue}}$$

Why we set the ionization

- ◆ If most O were O3+ the process
- ◆ $O3+ + e \rightarrow O2+ + hn$
- ◆ Would be fast, and would make O III recombination lines
- ◆ This can happen in nature, but it would confuse our homework problem

Vary density over extreme range

- ◆ Plot emissivity vs density over wide range to see how emissivity changes
- ◆ Recombination line, [O III] forbidden lines

Recombination lines

- ◆ $H^+ + e \rightarrow H^{0*} \rightarrow H^0 + \text{photons}$
- ◆ Critical densities of H I, He I, and He II optical lines are very high, $n > 1e15 \text{ cm}^{-3}$, so they are usually in LDL
- ◆ Emissivity goes as n^2 for $n < 10^{20} \text{ cm}^{-3}$
- ◆ Case B predictions
- ◆ H I, He I, He II are the strongest in UV/ Opt/ IR
- ◆ Second row (C,N, O, Ne) & Fe in X-ray

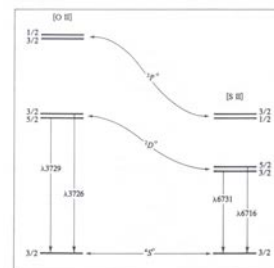
Forbidden lines

- ◆ [O III]
- ◆ $O^{++} + e \rightarrow O^{++*} + e \rightarrow O^{++} + e + \text{photons}$
- ◆ Critical densities of many forbidden lines $n \sim 1e3 \text{ cm}^{-3}$, so they can be in LDL or HDL
- ◆ Emissivity goes as n^2 or n

Compute spectrum of clouds with two very different densities

- ◆ $H_{den} = 4$
- ◆ $H_{den} = 14$
 - How will emission from these cloud compare?
 - How can we “trick” the model into having roughly the same emission?

Density indicators



AGN3 Fig 5.7

